

Periodic Table Basics

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

5

B

Boron


10.81

P = 5


N = 6

E = 5

Bohr Diagram:



Lewis structure:



Step 2: Determine the number of protons, neutrons, and electrons in each element.

Step 3: Create a Bohr diagram for each element.

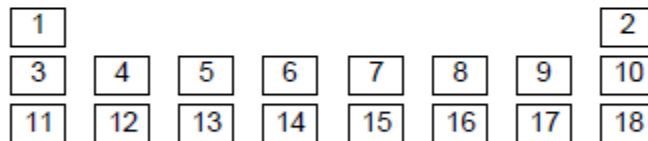
Step 4: Draw the Lewis Structure for each element.

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Make your own key if you don't have a color

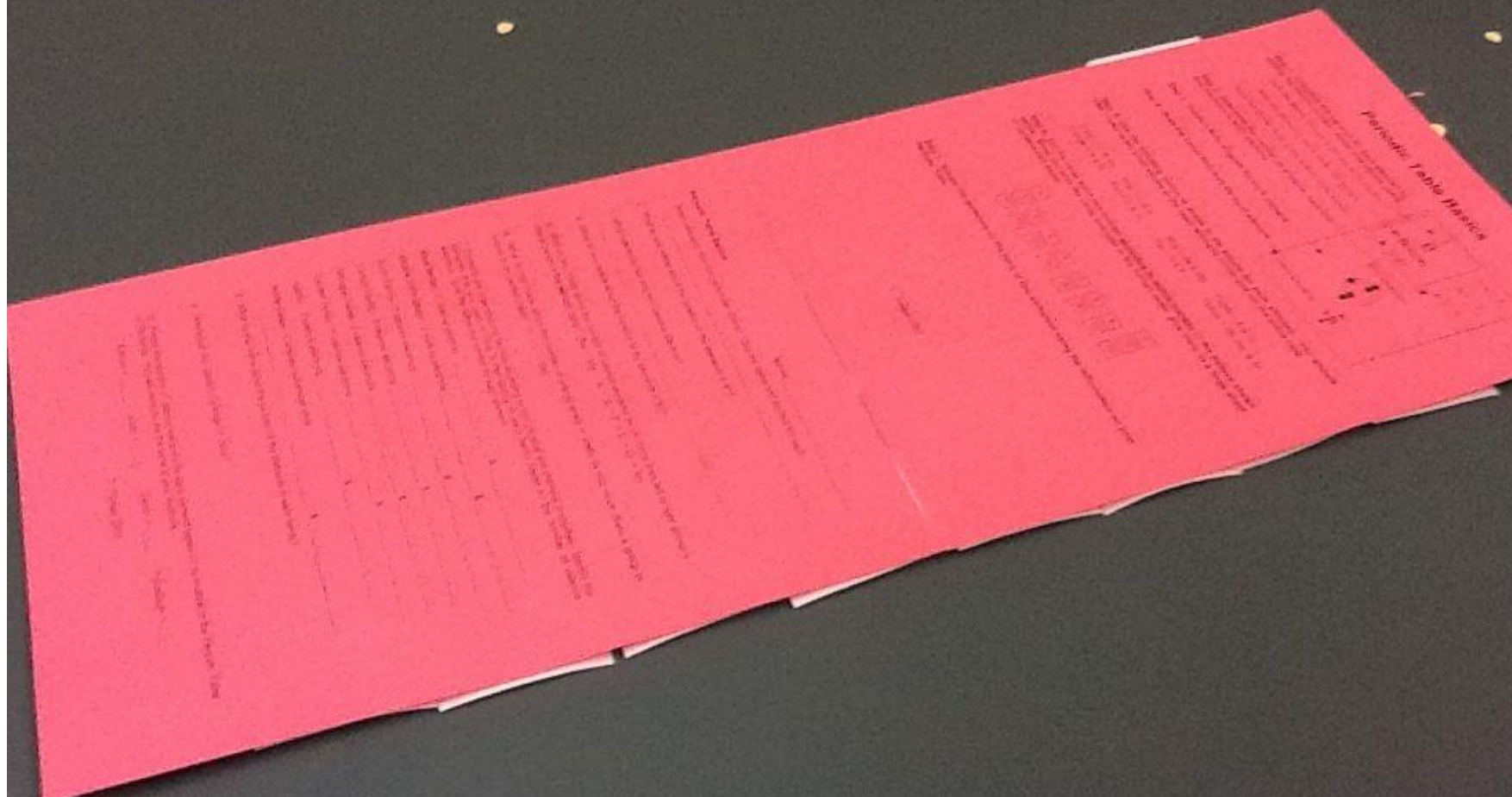
Green = Li & Na	Pink = O & S	Blue = Be & Mg	Purple = F & Cl
Orange = B & Al	Red = C & Si	Tan = N & P	Yellow = He, Ne, & Ar

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.

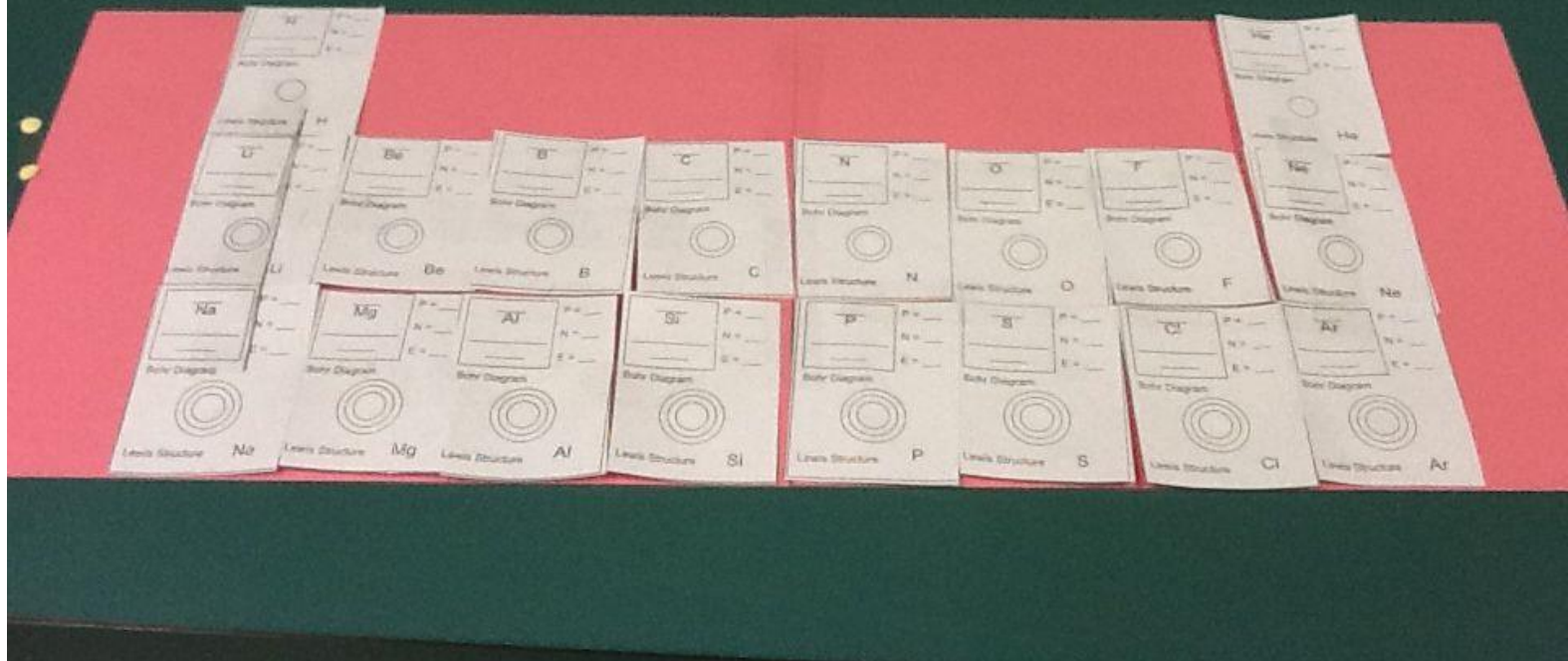


Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.

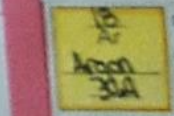
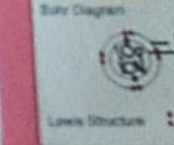
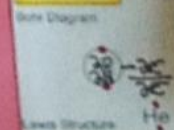
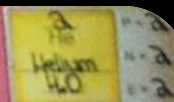
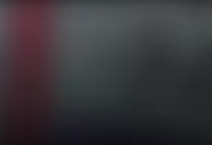
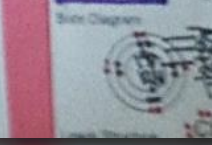
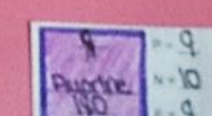
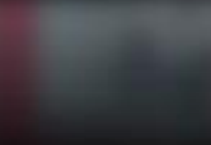
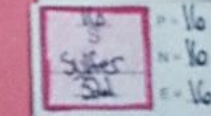
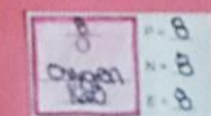
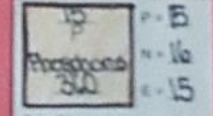
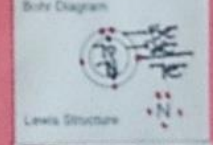
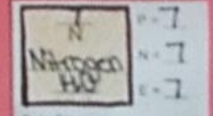
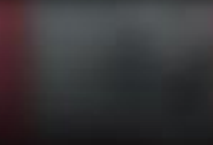
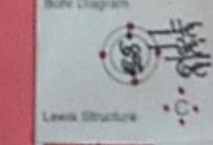
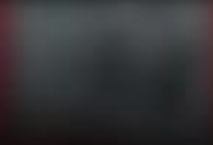
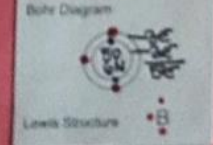
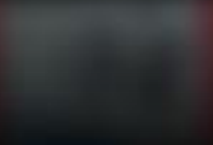
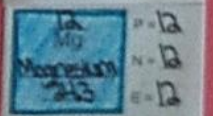
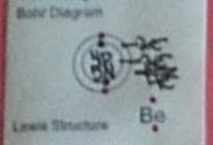
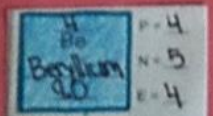
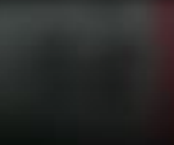
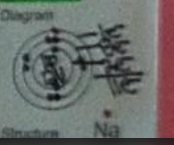
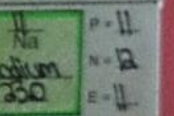
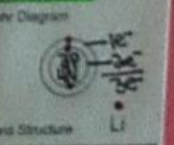
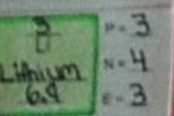
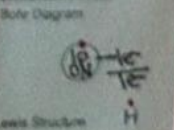
The Periodic Table Basics cover Sheet
And the QUESTIONS will be taped together like below



Then, the elements will be cut out and glued or taped onto the back
Once you have completed filling out the cards



Hydrogen
1.01
1
1
1



kelompok
7/10
nisa
(4/5 KMP?)

Lewis Structure



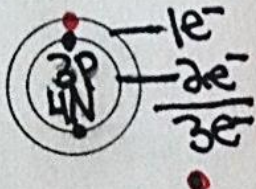
3
Li
Lithium
6.9

P = 3

N = 4

E = 3

Bohr Diagram



Lewis Structure



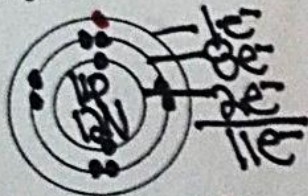
11
Na
Sodium
23.0

P = 11

N = 12

E = 11

Bohr Diagram



Lewis Structure



4
Be
Beryllium
9.0

P = 4

N = 5

E = 4

Bohr Diagram



Lewis Structure



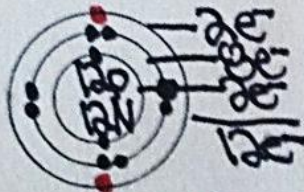
12
Mg
Magnesium
24.3

P = 12

N = 12

E = 12

Bohr Diagram



Lewis Structure



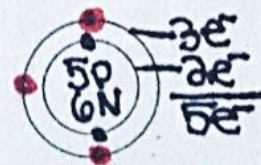
5
B
Boron
10.8

P = 5

N = 6

E = 5

Bohr Diagram



Lewis Structure



13
Al
Aluminum
27.0

P = 13

N = 14

E = 13

Bohr Diagram



Lewis Structure



EXAMPLE

$\frac{15}{P}$ Phosphorus 31.0	P = <u>15</u> N = <u>16</u> E = <u>15</u>
--------------------------------------	---

Bohr Diagram

Lewis Structure

Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

- | | | | |
|-----------------|--------------|----------------|-----------------------|
| Green = Li & Na | Pink = O & S | Blue = Be & Mg | Purple = F & Cl |
| Orange = B & Al | Red = C & Si | Tan = N & P | Yellow = He, Ne, & Ar |

Created with Docer

Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown.

