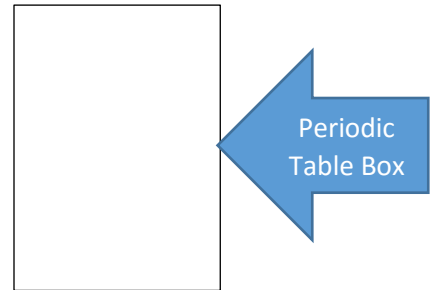


Isotope name: _____

Picture

Atomic #:	# of protons =
Symbol:	# of electrons =
Name:	# of neutrons =
Mass #(for this isotope):	



Isotope name: _____

Picture

:

Atomic #:	# of protons =
Symbol:	# of electrons =
Name:	# of neutrons =
Mass #(for this isotope):	

COLOR KEY for pictures: PROTONS NEUTRONS ELECTRONS

Explain why these two atoms are isotopes.

Which isotope is likely unstable and/or radioactive? Explain using the MASS # on the PERIODIC TABLE.

Mastery: Explain why the mass # is often a decimal. Make up an example or illustrate how this can be true

Last name: _____ First name: _____ Date: _____ Period: _____

ISOTOPES DRAWINGS and CHAPTER REVIEW

INSTRUCTIONS: Choose one of the following isotopes. Fill in the provided template with the requested information and draw each isotope.

Carbon-12 vs Carbon-14

Nitrogen-14 vs Nitrogen-16

Beryllium 9 vs Beryllium -10

5 4 3 2 1 0 Late

Checklist to Mastery (check ITEMS you completed WELL)

- Name** for each isotope written
- Atomic** number, **mass** number, **name**, and **symbol** written in **correct format**
- Correct number** of P, N, and e- (written and drawn)
- Labeled and color-coded** (P, N, e-)
- Explanation of why** atoms are **isotopes**
- Radioactivity** discussed
- Why the **DECIMAL on the Mass #?**
- Neat**
- Colorful**
- Creative**
- Chapter Review Completed**

Chapter Review: Questions 1-16 pp 296 -297

CIRCLE YOUR OPTION BELOW:

Regular	ADV / MASTERY
Pick 6 from 296	Pick 4 from 297

WRITE NEAT! Keep answers In Context when possible
