

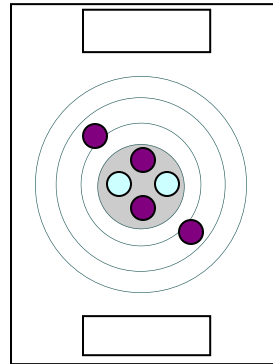
5 4 3 2 1 0 (inc)

FORMATIVE: ATOMS really add up!

Directions: Using your notes and your textbook to complete the following.

1. Label the atom with

- NUCLEUS
- PROTONS (+)
- NEUTRONS (+/-)
- ELECTRONS (-)
- ATOMIC MASS
- ATOMIC NUMBER



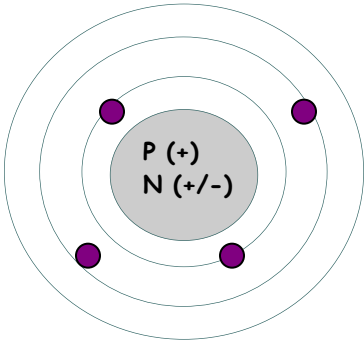
2. The **atomic number** equals the number of _____ in the nucleus.
3. The **protons** have a _____ charge while the _____ have a **negative charge**.
4. Only the **protons** and the _____ have **measurable mass**.
5. Therefore, **atomic mass** = _____ + **neutrons**.
6. To find the **number of negative charges**, find the number of _____ charges.
because the number of _____ = the number of electrons.
7. **Number of neutrons** is equal to the **ATOMIC MASS** - _____.

<i>Element Name</i>	<i>Symbol</i>	<i>Atomic Number</i>	<i>Number Neutrons</i>	<i>Atomic Mass</i>
Iodine		53	74	
	H	1	0	
Zinc		30	35	
Gold		79	118	
Tungsten		74	110	
	Kr	36	48	
*	Co	27	32	
*Sulfur		16	16	
*Potassium		19	20	
*Arsenic		33	42	

8.

<i>Element</i>	<i>Symbol</i>	<i>Protons</i>	<i>Neutrons</i>	<i>Atomic Mass</i>	<i>Electrons</i>	<i>Atomic Number</i>
Carbon			6	12		
	K	19	20			19
*Oxygen		8		16		
*Sodium				23	11	

9.



Protons + Neutrons = Atomic Mass
 Atomic Mass - Protons = Neutrons
 Protons (+) = Electrons (-)

ATOMIC MASS

<i>Element</i>	<i>Symbol</i>	<i>Protons + Neutrons =</i>		<i>Atomic Mass</i>
	O	8	8	
Sodium		11		
	Al	13		27
*Scandium			24	45
*	Zn	30	35	65

<i>Element</i>	<i>Symbol</i>	<i>Atomic Number</i>	<i>Electrons</i>	<i>Protons</i>	<i>Neutrons</i>	<i>Atomic Mass</i>
Lithium	Li	3	3	3	4	7
	Be	4				9
Carbon		6			6	
		14				28
				5	6	
	F		9			19
*Phosphorus		15				31
*				16	16	
*			20		20	
*		33				75
*				40	51	